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| What will we be learning? Acid and Alkali | Why this? Why now? Previous units-- Metal and Non-metals, Earth structure and separating mixturesNext units- Chemical Energy, Types of reactions.Enquiry ProcessesAnalyse Patterns, Draw conclusions, Present data, Justify opinions, Collect data, Present data, Plan variables | Key Words:alkali, acid, neutralisation, salt, hydrochloric acid, sulphuric acid, corrosive, dilute, concentrated, hazard, irritant, indicator, antacid, neutralise, universal indicator, pH scale, chloride, sulphate. |
| What will we learn?To identify if a substance is an acid or an alkali from its pH number.Classify solutions as acidic, alkaline or neutral using indicator colours and pH values.State that in a neutralisation reaction a chemical called a salt and water is formed.State that in a neutralisation reaction a chemical called a salt and water is formed.Write a set of instructions to carry out the experimentMisconceptions in this topic* Some students may think that antacids are acidic substances and so called because they do not react with acids.
* Students will use the term strong/weak and concentrated/dilute to mean the same thing-. A more concentrated solution of a weak acid could, therefore, have a lower pH. An acid’s strength is a measure of how easily the acid dissociates (gives up its hydrogen
* Acids and alkalis are dangerous.’
* Students often confuse the concepts of neutralisation and dilution.
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| What opportunities are there for wider study?STE(A)M- <https://highcliffe.sharepoint.com/sites/LearnSTEM> |
| How will I be assessed?End of topic assessmentInvestigation markingPeer Assessment |